**Grade 8 – REVIEW – Test on the Periodic Table**

1. As you move from top to bottom down the periodic table the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (increases or decreases).
2. Where would you expect to find the smallest atoms of elements in the periodic table?
3. There are only \_\_\_\_ nonmetal elements.
4. The group of elements also referred to as semimetals is the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
5. The ***Ag*** in the periodic table stands for \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
6. The ***Au*** in the periodic table stands for \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
7. The ***Hg*** in the periodic table stands for \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
8. The ***K*** in the periodic table stands for \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
9. The rows of the periodic table are called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. The scientist credited with devising the first periodic table similar to the one we use today was \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
11. The periodic law is just a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
12. The periodic table is a table for the .
13. The primary difference between the modern periodic table and Mendeleev's periodic table is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
14. As you move from left to right across the periodic table the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ increases.
15. State an example of an alkali metal - \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
16. How many natural elements are there on the periodic table? \_\_\_\_\_\_\_
17. The Carbon Group is composed of the following elements:
18. The Oxygen Group is composed of the following elements:
19. Substances called *\_\_\_\_\_\_\_\_\_\_\_\_\_* can be found in the transition elements, they speed up chemical reactions.
20. Iron, cobalt and nickel form the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
21. *Alkali metals* are in group \_\_\_\_\_\_\_\_\_\_\_\_\_on the Periodic table of Elements
22. The vast majority of elements in the periodic table are classified as \_\_\_\_\_\_\_\_\_\_\_\_
23. The periodic table is composed of \_\_\_\_\_\_\_\_\_\_\_periods and \_\_\_\_\_\_\_\_\_\_\_\_groups.
24. Each family in the periodic table has its own characteristics properties based on the number of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
25. List 3 properties of nonmetals: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
26. List the different types of metals: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
27. Periodic law states that elements show a …
28. How is the current periodic table organized?
29. What is the same within each group on the Periodic Table?
30. Why don't noble gases normally form chemical bonds?
31. Which property of metals means that it can be drawn into wires?

1. What can group numbers on the periodic table help you determine?

**Grade 8 – REVIEW ANSWER KEY – Test on the Periodic Table**

1. As you move from top to bottom down the periodic table the \_\_**NUMBER OF ENERGY LEVELS\_\_ increases**.
2. Where would you expect to find the smallest atoms of elements in the periodic table?

# PERIOD 1 – UPPER LEFT CORNER

1. There are only **17**\_ nonmetal elements.
2. The group of elements also referred to as semimetals is the \_\_\_**METALLOIDS**\_\_.
3. The ***Ag*** in the periodic table stands for \_\_\_\_**SILVER**\_\_\_.
4. The ***Au*** in the periodic table stands for \_\_\_\_**GOLD**\_\_\_\_\_\_.
5. The ***Hg*** in the periodic table stands for \_\_\_\_\_**MERCURY**\_\_\_.
6. The ***K*** in the periodic table stands for \_\_\_\_**POTASSIUM**\_\_\_.
7. The rows of the periodic table are called \_\_**PERIODS**\_\_
8. The scientist credited with devising the first periodic table similar to the one we use today was \_\_\_\_**MENDELEEV**\_\_\_
9. The periodic law is just a \_\_\_**THEORY**\_\_
10. The periodic table is a table for the **ELEMENTS**.
11. The primary difference between the modern periodic table and Mendeleev's periodic table is \_\_\_**MODERN DAY ARRANGED BY ATOMIC NUMBER --- MENDELEEV WAS ARRANGED BY ATOMIC MASS**\_\_\_\_\_\_\_
12. As you move from left to right across the periodic table the \_\_**NUMBER OF VALENCE ELECTRONS**\_ increases.
13. State an example of an alkali metal - \_(any one of these is correct: \_**LITHIUM; SODIUM; POTASSIUM; RUBIDIUM; CESIUM; FRANCIUM\_\_**
14. How many natural elements are there on the periodic table? \_**92**\_
15. The Carbon Group is composed of the following elements:

# Carbon – silicon – germanium – tin - lead

1. The Oxygen Group is composed of the following elements:

# Oxygen - sulfur – selenium – tellurium - polonium

1. Substances called *\_***CATALYSTS***\_\_* can be found in the transition elements, they speed up chemical reactions.
2. Iron, cobalt and nickel form the \_**IRON TRIAD**\_\_\_
3. *Alkali metals* are in group \_\_**1** \_on the Periodic table of Elements
4. The vast majority of elements in the periodic table are classified as \_\_**METALS**\_
5. The periodic table is composed of \_**\_7**\_\_periods and \_\_\_**\_18**\_\_groups.
6. Each family in the periodic table has its own characteristics properties based on the number of \_**VALENCE ELECTRONS**\_\_
7. List 3 properties of nonmetals: \_\_\_BRITTLE\_\_, \_\_\_\_**POOR CONDUCTOR OF HEAT**\_\_ and \_\_\_\_ **POOR CONDUCTOR OF ELECTRICITY**\_\_\_\_\_\_
8. List the different types of metals: \_\_**ACTIVE METALS**\_\_**ALKALI**\_\_, \_\_\_**ALKALI EARTH**\_\_ and \_\_**RARE EARTHS**\_\_
9. Periodic law states that elements show a … **REPETITION OF THEIR PROPERTIES WHEN ARRANGED BY INCREASING ATOMIC NUMBER**\_\_\_
10. How is the current periodic table organized? **BY ATOMIC NUMBER**
11. What is the same within each group on the Periodic Table?  **THE NUMBER OF VALENCE ELECTRONS**
12. Why don't noble gases normally form chemical bonds? **THEY ALL HAVE FILLED OUTER ENERGY LEVELS**
13. Which property of metals means that it can be drawn into wires?  **DUCTILE**
14. What can group numbers on the periodic table help you determine? **THE NUMBER OF VALENCE ELECTRONS**